

The Infrared Spectra and Optical Constants of
Elusive Amorphous Methane

Short title: IR spectrum and optical constants of amorphous methane

Perry A. Gerakines¹

and

Reggie L. Hudson^{1,2}

Submitted to *The Astrophysical Journal Letters*

¹ Astrochemistry Laboratory, NASA Goddard Space Flight Center, Greenbelt, Maryland,
20771 USA

² E-mail: Reggie.Hudson@NASA.gov

ABSTRACT

New and accurate laboratory results are reported for amorphous methane (CH₄) ice near 10 K for the study of the interstellar medium (ISM) and the outer Solar System. Near- and mid-infrared (IR) data, including spectra, band strengths, absorption coefficients, and optical constants, are presented for the first time for this seldom-studied amorphous solid. The apparent IR band strength near 1300 cm⁻¹ (7.69 μm) for amorphous CH₄ is found to be about 33% higher than the value long used by IR astronomers to convert spectral observations of interstellar CH₄ into CH₄ abundances. Although CH₄ is most likely to be found in an amorphous phase in the ISM, a comparison of results from various laboratory groups shows that the earlier CH₄ band strength at 1300 cm⁻¹ (7.69 μm) was derived from IR spectra of ices that were either partially or entirely crystalline CH₄. Applications of the new amorphous-CH₄ results are discussed, and all optical constants are made available in electronic form.

Key Words: Astrochemistry; infrared: ISM; ISM: abundances; ISM: molecules; planets and satellites: surfaces; molecular data

1. INTRODUCTION

At present we are engaged in an extensive program to record infrared (IR) reference spectra of solids at low temperatures with which to determine complex indices of refraction ($n - ik$). Such results are needed for studies of both interstellar and planetary chemistry, mainly for determinations of abundances of ices in the outer Solar System and the interstellar medium (ISM). To date we have examined the simplest members of the alkyne and alkene classes of hydrocarbons, acetylene and ethylene, and have found significant and unexpected problems with the literature on both molecules to the extent that our new work replaces much of the old (Hudson et al. 2014a, 2014b). Recently we have turned our attention to CH₄, methane, the simplest alkane, but here too we have encountered problems with the literature results and these have motivated this Letter.

Methane is one of the few molecules that form ices of interest to both interstellar and planetary astrochemists. It can, in principle, exist in three phases under vacuum or near-vacuum conditions. The high-temperature phase is a plastically-crystalline solid with a degree of rotational disorder lacking in the low-temperature, more-ordered crystalline form, the reversible transition between the two being at 20.4 K (Clusius 1929). The third CH₄ phase is an amorphous one formed by the slow condensation of methane from the gas phase at temperatures on the order of 20 K and lower. Although limited sets of optical constants for the two crystalline phases have been published (Pearl et al. 1991; Hudgins et al. 1993), we have not found optical constants for CH₄ ices that are unequivocally amorphous. Further, we have discovered significant

questions about the amorphous character of essentially all CH₄ ice samples whose spectra appear in the literature.

As an example of these difficulties, Figure 1(a-e) shows the 1300 cm⁻¹ (ν₄) IR feature of CH₄ deposited at ~15 K as published over about 20 years by five independent sets of researchers. The degree of reproducibility is reassuring as in each case the ν₄ absorption is seen to have two distinct components with about the same relative intensities. However, we have found (*vide infra*) that warming a CH₄-ice sample that possesses these same spectral characteristics causes the two peaks to coalesce, and that recooling such ices to ~15 K then regenerates the original spectra of Figure 1(a-e), an unusual reversible behavior for amorphous organic compounds, but one in agreement with the behavior of methane's two crystalline phases (Savitsky & Hornig 1962). We conclude that despite the reported temperatures of CH₄ condensation, the spectra of Figure 1(a-e) are *not for an amorphous ice, but for a crystalline CH₄ sample* formed on deposition.

Knowing the phase of a laboratory sample is important as many ices in the interstellar medium and the outer Solar System are expected to be amorphous, and so efforts to quantitatively analyze them for CH₄ will be severely hindered without an appropriate laboratory reference spectrum and band strengths. Moreover, a proper, well-characterized set of IR features is critical for any accurate calculation of optical constants. This Letter addresses such problems by presenting IR spectra of amorphous CH₄ along with associated band strengths, absorption coefficients, and optical constants, apparently the first for amorphous CH₄ in the literature.

2. EXPERIMENTAL METHODS

Descriptions of our experimental setup, ice preparation method, and spectral measurements have been published, so only a summary is needed here. See, for example, Hudson & Moore (2004) and Moore et al. (2010). Briefly, CH₄ (Matheson Research Grade) in a vacuum line was flowed through a controlled leak valve onto a pre-cooled (8 - 30 K) KBr substrate inside a stainless steel high-vacuum chamber ($\sim 10^{-7}$ torr). The sample gas was condensed at $\sim 1.5 \times 10^{17}$ molecules min⁻¹ to an ice thickness of 0.5 - 5 μm as determined by a laser (670 nm) interference fringe-counting system. After deposition, transmission spectra (0.5-cm⁻¹ resolution, 100 co-added scans) of the ice were recorded *in situ* with a Perkin-Elmer Spectrum GX Fourier-transform IR spectrometer equipped with a room-temperature DTGS detector. After the initial spectrum was recorded, the sample was warmed or cooled in increments of a few degrees with temperatures being measured (precision ~ 0.2 K) with a silicon diode inserted into the metal frame holding the KBr substrate. Spectra were recorded to document any changes in band position, intensity, and shape. Methane ices were studied at 8 - 40 K, and all samples rapidly sublimed at higher temperatures. The great majority of measurements were made in the 8 - 12 K range, and for simplicity we refer to these simply as at 10 K. As an extra check on our temperature calibration, we recorded IR spectra of crystalline CD₄ at 10 - 30 K, and observed phase changes to occur at 22 and 27 K, as expected (Clusius et al. 1937; Kinsel & Trikha 1994). Residual H₂O in the vacuum line and the CH₄ source was the only contaminant detected, and could be seen in trace amounts (≤ 0.01 absorbance units) in the 3700 - 3500 and 1600 cm⁻¹ regions of our spectra, far from the IR peaks of interest.

Methane's IR spectra were recorded from 5000 to 400 cm^{-1} (2.0 to 25 μm), with special attention being given to the molecule's ν_4 region (1300 cm^{-1} , 7.69 μm) and the near-IR features at 4600 - 4100 cm^{-1} (2.17 - 2.44 μm). Apparent band strengths (A') and apparent absorption coefficients (α') were determined from appropriate Beer's law plots as described in Hudson et al. (2014a) and references therein.

In calculating ice thicknesses, absorption coefficients, and band strengths of amorphous CH_4 , values of the ice density (ρ) and a reference value of the refractive index (n) were needed, both near 10 K. For the density we adopted $\rho = 0.47 \text{ g cm}^{-3}$ (Satorre et al., 2008), which gives a number density of $\rho_N = 1.77 \times 10^{22} \text{ molecules cm}^{-3}$. Using the two-laser method of our previous work with hydrocarbons (Hudson et al. 2014a, 2014b) we measured n of amorphous CH_4 five times at ~ 10 K and got an average value of $n = 1.28$ with a standard error of ± 0.008 . Our n is essentially identical to the average of the two values of Satorre et al. (2008, Figure 3) for a similar temperature and wavelength.

3. RESULTS

There are many similarities among the CH_4 spectra in (a) - (e) of Figure 1, but in our hands a different result was obtained. Figure 1(f) is the IR spectrum of CH_4 -ice made in our laboratory near 10 K. On either warming such samples to ~ 30 K or on depositing CH_4 directly at ~ 30 K, a sharp symmetric peak as in Figure 1(g) was observed, which subsequent recooling to 10 K changed into two well-resolved features of Figure 1(h) that matched those shown in Figure 1(a - e). The conversion between the single sharp feature near 30 K and the doublet near 10 K was completely reversible,

as shown in Figure 1(g - j). Additional work showed that either depositing CH₄ above ~20 K and then cooling or depositing about ten times faster at 10 K also produced the double-peaked absorption. Very slow coolings revealed that the transition to the double-peaked pattern consistently occurred at 20 - 21 K. Two confirmations of our results came through independent spectral measurements of CH₄ deposited at 8.8 K on an aluminum substrate and in methane work by one of us on a similar experimental set-up in a different laboratory (Gerakines et al. 2005). We emphasize that the conversion from the spectrum of Figure 1(f) to that of 1(g) was irreversible, as expected for an amorphous-to-crystalline phase change. Combining all of these results, we assign the spectrum of Figure 1(f) to amorphous CH₄. The spectra in 1(g) and 1(i) are of phase-I crystalline CH₄, and traces 1(h) and 1(j) are for phase-II crystalline CH₄.

Figure 2(a) shows enlargements of three regions of the spectrum of amorphous-CH₄, for a deposition at ~10 K. We combined such spectra with a Kramers-Kronig routine to calculate the optical constants n and k of amorphous CH₄, which also are shown in Figure 2. Spectra and optical constants are available in electronic form on our group's website (<http://science.gsfc.nasa.gov/691/cosmicice/constants.html>) for the 5000 - 400 cm⁻¹ range studied.

4. DISCUSSION

The interpretation of our results is straightforward. On slowly depositing methane gas near 10 K, amorphous CH₄-ice was formed, which then crystallized irreversibly into methane's high-temperature solid phase (phase I) on warming to 30 K. Recooling generated the low-temperature crystalline phase (phase II) of CH₄, with a transition near

20.4 K (Clusius 1929; Savitsky & Hornig 1962). The reversible cycling between spectral band shapes shown in Figure 1(g - j) can be taken as the *sine qua non* for crystalline CH₄, but we are unaware of a similar figure in the astrochemical literature that shows this high degree of reversibility. The pronounced asymmetry of the 1300-cm⁻¹ feature is similar to that of other intense small-wavenumber (long-wavelength) absorptions in amorphous hydrocarbons, such as acetylene (Hudson et al. 2014a) and ethane (Hudson et al. 2014b). See Keefe et al. (2012) for a more-general discussion of such shapes.

4.1. Optical Constants and Band Strengths

To the best of our knowledge, Figure 2 is the first publication of the IR features of amorphous CH₄. Similarly, the optical constants of Figure 2 are the first of their kind. These spectra and optical constants can be valuable in studies of low-temperature CH₄ ices in astronomical environments, but even more so if combined with measures of band strengths. As in our recent papers on C₂H₂ and C₂H₄, we distinguish between apparent and absolute measures of intensity (Hudson et al. 2014a, 2014b). Apparent intensities, either absorption coefficients (α) of spectral peaks or integrated intensities (A') of IR bands, are found by graphing either a peak height or band area as a function of ice thickness and evaluating the slope of the resulting line, similar to work with standard Beer's law plots (further details in Hudson et al. 2014a). For the ices used in nearly all laboratory studies of this type, α and A' include the effects of reflections and optical interference in ways that are difficult to predict. In contrast, absolute absorption

coefficients (α) of spectral peaks and absolute integrated intensities (A) of IR bands are found from optical constant k as follows:

$$\alpha(\tilde{\nu}) = 4\pi\tilde{\nu}k(\tilde{\nu}) \quad (1)$$

$$A = \frac{1}{\rho_N} \int_{band} \alpha(\tilde{\nu}) d\tilde{\nu} \quad (2)$$

where $\tilde{\nu}$ is wavenumber, almost always in units of cm^{-1} , and ρ_N is the number density of molecules in the ice, here in units of molecules cm^{-3} . This designation between apparent and absolute measures of intensity has long been made in the molecular spectroscopy literature (e.g., Maeda & Schatz 1961; Wexler 1967; Scrocco et al. 1972), but has not always been recognized in astrochemical papers.

Table 1 lists α' , α , A , and A' for several regions of the IR spectrum of amorphous CH_4 . All published values we have found for CH_4 ices at 10 - 20 K are for samples that are either partially or wholly crystalline, and so it is not surprising that they differ somewhat from our numbers in Table 1. (Note also that most of the quantities designated by A in the astrochemical literature on ices are, in fact, A' .) A full analysis of the uncertainties for all of the band strengths and absorption coefficients in Table 1 is in progress and results will be given in a follow-on paper. However, at present we can report that for both ν_3 and ν_4 , the features of most interest to interstellar astronomers, we find $A \approx A'$ with a relative uncertainty of about 1% for $A(\nu_3)$, $A'(\nu_3)$, $A(\nu_4)$, and $A'(\nu_4)$.

4.2. Implications for Laboratory Measurements on CH_4 -Ice

Part of the value of our new results is that they can help evaluate the phase of frozen CH_4 when such information is otherwise not available. For the CH_4 feature of

Figure 1, each of the spectral traces in (a) - (e) is now seen to be either that of phase-II crystalline CH₄-ice or perhaps the crystalline material contaminated with some amount of amorphous solid, despite the fact that all five samples were reported as being made at 10 - 20 K. Key to this assessment is our mid-IR result for amorphous CH₄ and knowledge of the double-peaked feature at 1300 cm⁻¹ for phase-II crystalline CH₄. Turning to the near-IR region, the paper of Bernstein et al. (2006) on H₂O + CH₄ mixtures included a near-IR spectrum of pure CH₄ deposited at 15 K. A distinct feature at 4314 cm⁻¹ in their Figure 2 shows that their sample was at least partially crystalline, as no such absorption is present in amorphous CH₄, but it is present in phase-II crystalline CH₄ (Baciocco et al. 1987; Khanna & Ngoh 1990). The same peak in Figure 4 of Gálvez et al. (2009) also betrays some crystallinity for their CH₄ ice sample, which was deposited at 14 K.

In some of these cases the results derived from such spectra will hardly depend on the initial CH₄ phase, such as if product identification is the goal in a photo- or radiation chemistry experiment. However, in other cases the failure to accurately recognize the phase of methane ice can lead to severe difficulties, such as in kinetic studies where a correct initial abundance may be needed to work out a mass balance, but an inappropriate band strength is adopted, or in trying to establish trends and draw conclusions from widths and positions of methane spectral bands.

An important point here is that neither the temperature at which methane is condensed nor the temperature at which a spectrum is recorded is a universal and reliable indicator of the phase of a CH₄ ice sample. Our experiments showed that either rapid deposition of CH₄ at ~10 K or deposition above ~20 K, followed by cooling,

produced the double-peaked 1300-cm^{-1} features of phase-II crystalline CH_4 . This suggests that the reason the published spectra of Figure 1(a-e) are not of pure amorphous CH_4 is because of (a) too high a deposition rate, leading to excessive energy release on gas condensation, (b) too poor a thermal conductivity for the substrate used, (c) too low of a cryostat cooling efficiency, (d) the location of the temperature sensor, or (e) some combination of these factors. We assume that these same spectra faithfully represent what was measured in each laboratory, but they should not be taken *a priori* as the spectra of amorphous CH_4 in the interstellar medium or the outer Solar System.

Many other interesting and relevant papers can be found in the CH_4 -ice literature. Older work by Pipes et al. (1978) is still valuable, but the intense ν_3 and ν_4 fundamental bands of their methane spectra appear saturated, prohibiting an assessment of the phase of their samples. Also, their use of a 20-K deposition temperature is of concern, being close to the 20.4 K phase-transition temperature. An early paper by d'Hendecourt & Allamandola (1986) was influential in describing the calculations needed and spectra expected for laboratory work on astronomical ices. Their 10-K methane ice appears to have been crystalline, not amorphous, from the positions of the two components reported near 1300 cm^{-1} .

Among the more-recent papers, that of Mejía et al. (2013) is interesting for many reasons, not least of which is that the lower spectrum in their Figure 1 may well be that of amorphous CH_4 . However, the small size shown for key spectral features hinders an independent evaluation and the figure's lack of a vertical numerical scale prohibits the extraction of quantitative information, such as α' values, which could help determine the

ice's phase. It also is curious that certain features of the spectrum after ion irradiation of the sample are said to be consistent with those of phase-II crystalline CH₄, whereas irradiation typically amorphizes ices rapidly at 15 K. Similarly, it is difficult to know what is meant when a methane spectrum is described as having "some amorphous features" (de Barros et al. 2011) when the present paper is the first to present the spectrum of amorphous CH₄.

In many cases, published IR transmission spectra are on too small a scale to permit independent verification, examination, and reproduction of the authors' work, and similar comments apply to IR spectra of CH₄ recorded with reflectance-absorbance methods (e.g., Moore & Hudson 2003; Bennett et al. 2006). This is not the case with the long paper of Hudgins et al. (1993). We initially were surprised that the CH₄ spectra shown there were said to be for an ice with a thickness of 0.03 μm, but the authors' inclusion of useful enlargements of key spectral regions, and detailed descriptions of experiments, revealed that the sample's actual thickness was 0.3 μm.

4.3. Implications for Astronomical Observations of CH₄-Ice

An immediate implication of our work is the need to reassess the CH₄ band strengths and optical constants used by astronomical observers. The strength of the methane band at 1300 cm⁻¹ (7.69 μm), as reported by Hudgins et al. (1993) and later modified by Boogert et al. (1997), is 7.3×10^{-18} cm molec⁻¹. This value and scaled versions for various ice densities have been used in analyzing both ISO (Gibb et al. 2004) and Spitzer (Öberg et al. 2008) infrared observations. The spectrum on which this band strength is based is shown in Figure 1(d), taken from Hudgins et al. (1993).

Our conclusion is that until now all of the methane IR spectra, band strengths, and optical constants available for interstellar studies in this spectral region are for crystalline CH₄, presumably the low-temperature phase (phase II) of the solid. In contrast, our Table 1 shows that the intensity of the corresponding IR band of amorphous CH₄ is $A'(1300\text{ cm}^{-1}) = 9.71 \times 10^{-18}\text{ cm molec}^{-1}$, which is about 33% higher than the crystalline-CH₄ value.

Although our measurements were made with pure CH₄, they have a significant bearing on the derived abundances of interstellar CH₄ mixed with H₂O-ice. The methane absolute intensities often used by IR astronomers are based on those of pure CH₄, but scaled to the IR band area of a laboratory sample of CH₄ in a mixed molecular ice of known composition (e.g., Kerkhof et al. 1999). Moreover, the reference CH₄ band intensity for this scaling typically is taken from data on crystalline CH₄ (d'Hendecourt & Allamandola, 1986; Hudgins et al. 1993), as opposed to the amorphous CH₄ results presented here. A similar scaling for the intensity of the methane feature near 1300 cm⁻¹ (7.69 μm), using the results for amorphous CH₄ in our Table 1, would give a larger band strength for CH₄ in H₂O-ice. Our band strengths can, of course, be re-scaled using appropriate densities to estimate A' for amorphous CH₄ in mixtures dominated by others solids (e.g., H₂O, CO₂). See Boogert et al. (1997) for examples.

Another implication of our results is that the spectral models of cold planetary bodies, primarily Kuiper Belt objects, that rely on the optical constants of CH₄ in the near-IR might need re-evaluation. Our optical constants should be helpful in determining the physical state of cold objects on the edge of the Solar System and unraveling their hydrocarbon chemistry (Brown et al. 2007; Cruikshank et al. 2015).

The main spectral data available until now for CH₄ at these wavelengths are the absorption coefficients of Quirico & Schmitt (1997), which are for crystalline CH₄. Whether, for example, methane mixes with nitrogen to form an amorphous solid at ~30 K or whether the two materials segregate to form crystalline solids or whether another scenario applies is unknown, partly due to the lack of near-IR optical constants for amorphous CH₄ with which to interpret and synthesize spectra. Our results fill this gap.

5. CONCLUSION

We have presented for the first time details of the near- and mid-IR spectra of surprisingly-elusive amorphous CH₄-ice along with IR optical constants, absorption coefficients, and band strengths. Both apparent and absolute versions of α and A have been tabulated and all results made available online. The apparent band strength near 10 K of the amorphous-CH₄ fundamental near 1300 cm⁻¹ (7.69 μ m) has been found to be about 33% larger than the value for crystalline CH₄. With IR data on both the amorphous and crystalline phases of solid CH₄ finally available, it should now be possible to compute accurate CH₄ abundances from astronomical spectra and to better correlate them with temperature, the presence of other molecules, and so forth. Additional work is in progress.

NASA funding through the Outer Planets Research and Cassini Data Analysis programs is acknowledged. Both authors received partial support from the NASA Astrobiology Institute through the Goddard Center for Astrobiology. The experimental assistance of Marla Moore, Mark Loeffler, and Robert Ferrante is acknowledged.

REFERENCES

- Alsindi, W. Z., Gardner, D. O., van Dishoeck, E. F., & Fraser, H. J. 2003, *Chem. Phys. Lett.*, 378, 178
- Baciocco, G., Calvani, P., & Cunsolo, S. 1987, *Intl. J. Infrared Millimeter Waves*, 8, 923
- Bennett, C. J., Jamieson, C. S., Osamura, Y., & Kaiser, R. I. 2006, *ApJ*, 653, 792
- Bernstein, M. P., Cruikshank, D. P., & Sandford, S. A. 2006, *Icarus*, 181, 302
- Boogert, A. C. A., Schutte, W. A., Helmich, F. P., Tielens, A. G. G. M., & Wooden, D. H. 1997, *A&A*, 317, 929
- Brown, M. E., Barkume, K. M., Blake, G. A., Schaller, E. L., Rabinowitz, D. L., Roe, H. G., & Trujillo, C. A. 2007, *ApJ*, 133, 284
- Brunetto, R., Caniglia, G., Baratta, G. A., & Palumbo, M. E., 2008, *ApJ*, 686, 1480
- Clusius, K. 1929, *Zeitschrift für Phys. Chem.*, B3, 41
- Clusius, K., Popp, L., & Frank, A. 1937, *Physica*, 4, 1105
- Cruikshank, D. P., Grundy, W. M., DeMeo, F. E., Buie, M. W., Binzel, R. P., Jennings, D. E., Olkin, C. B., Parker, J. W., Reuter, D. C., Spencer, J. R., Stern, S. A., Young, L. A., & Weaver, H. A. 2015, *Icarus*, 246, 82
- de Barros, A. L. F., Bordalo, V., Seperuelo Duarte, E., da Silveira, Domaracka, A., Rothard, H. & Boduch, P. 2011, *A&A*, 531, A160
- d'Hendecourt, L. B., & Allamandola, L. J. 1986, *A&AS*, 64, 453
- Gálvez, Ó., Maté, B., Herrero, V. J., & Escribano, R. 2009, *ApJ.*, 703, 2101
- Gerakines, P. A., Schutte, W. A., & Ehrenfreund, P. 1996, *A&A*, 312, 289. Spectra: <http://www.strw.leidenuniv.nl/lab/databases/ice-analogues-schutte/database.html>
- Gerakines, P. A., Bray, J. J., Davis, A., & Richey, C. R. 2005, *ApJ*, 620, 1140

Gibb, E. L., Whittet, D. C. B., Boogert, A. C. A., & Tielens, A. G. G. M. 2004, *ApJS*, 151, 35

Hudgins, D. M., Sandford, S. A., Allamandola, L. J., & Tielens, A. G. G. M. 1993, *ApJSS*, 86, 713

Hudson, R. L., & Moore, M. H. 2004, *Icarus*, 172, 466

Hudson, R. L., Ferrante, R. F., & Moore, M. H. 2014a, *Icarus*, 228, 276

Hudson, R. L., Gerakines, P. A., & Moore, M. H. 2014b, *Icarus*, 243, 148

Kansal, O., & Trikha, S. K. 1994, *Aust. J. Phys.*, 47, 97

Keefe, C. D., Wilcox, T., & Campbell, E. 2012, *J. Molec. Struct.*, 1009, 111

Kerkhof, O., Schutte, W. A., & Ehrenfreund, P. 1999, *A&A*, 346, 990

Khanna, R. K., & Ngoh, M. 1990, *Spectrochim. Acta*, 46, 1057

Maeda, S., & Schatz, P. N. 1961, *J. Chem. Phys.* 35, 1617

Mejía, C. F., de Barros, A. L. F., Bordalo, V., da Silveira, E. F., Boduch, P., Domaracka, A., & Rothard, H. 2013, *MNRAS*, 433, 2368

Moore, M. H., & Hudson, R. L. 2003, *Icarus*, 161, 486

Moore, M. H., Ferrante, R. F., Moore, W. J., & Hudson, R. L. 2010, *ApJS*, 191, 96

Mulas, G., Baratta, G. A., Palumbo, M. E., & Strazzulla, G. 1998, *A&A*, 333, 1025

Öberg, K. I., Boogert, A. C. A., Pontoppidan, K. M., Blake, G. A., Evans, N. J., Lahuis, F., & van Dishoeck, E. F. 2008, *ApJ*, 678, 1032

Pearl, J., Ngoh, M. Ospina, M., & Khanna, R. 1991, *JGR - Planets*, 96, 17477

Pipes, J. G., Roux, J. A., Smith, A. M., & Scott, H. E. 1978, *AIAA J.*, 16, 984

Quirico, E., & Schmitt, B., 1997, *Icarus*, 127, 354

Satorre, M. Á., Domingo, M., Millán, C., Luna, R., Vilaplana, R., & Santonja, C. 2008,
Planet. Space Sci., 56, 1748

Savitsky, G. B., & Hornig, D. F. 1962, J. Chem. Phys., 36, 2634

Scrocco, M., Giuliani, R., & Costarelli, C. 1972, Spectrochim. Acta, 28A, 761

Wexler, A. 1967, Appl. Spectrosc. Rev. 1, 29

Figure Captions

Figure 1. Infrared spectra of the ν_4 band of CH_4 . Left column: CH_4 deposited at (a) 12.5 K, Mulas et al. (1998), (b) 10 K, Gerakines et al. (1996), (c) 15 K, Alsindi et al. (2003), (d) 10 K, Hudgins et al. (1993), and (e) 15 K, de Barros et al. (2011). Right column: CH_4 deposited in the present work at (f) ~ 10 K, (g) then warmed to 30 K and held there about 8 hours, (h) recooled to 10 K, (i) rewarmed to 30 K, and (j) recooled to 10 K.

Figure 2. Enlargements of selected sections of the near- and mid-infrared spectrum and associated optical constants n and k of amorphous CH_4 at ~ 10 K. The CH_4 sample whose spectrum is shown had a thickness of $0.5 \mu\text{m}$.

Table 1Peak positions, absorption coefficients, and band strengths for amorphous CH₄ at 8 K

Peak Position ^a cm ⁻¹ (μm)	α' cm ⁻¹	Peak Position ^b cm ⁻¹ (μm)	α cm ⁻¹	Integration range cm ⁻¹	A' cm molecule ⁻¹	A cm molecule ⁻¹
4529 (2.208)	107	4529 (2.208)	114	4559-4496	8.76 × 10 ⁻²⁰	9.37 × 10 ⁻²⁰
4301 (2.325)	879	4301 (2.325)	893	4428-4246	8.15 × 10 ⁻¹⁹	8.55 × 10 ⁻¹⁹
4203 (2.379)	514	4202 (2.380)	528	4240-4167	3.89 × 10 ⁻¹⁹	3.97 × 10 ⁻¹⁹
4115 (2.430)	14	4115 (2.430)	16	4144-4094	1.29 × 10 ⁻²⁰	1.75 × 10 ⁻²⁰
3846 (2.600)	86	3846 (2.600)	95	3853-3836	3.24 × 10 ⁻²⁰	4.16 × 10 ⁻²⁰
3010 (3.322)	21,860	3011 (3.321)	23,810	3160-2930	1.40 × 10 ⁻¹⁷	1.41 × 10 ⁻¹⁷
2814 (3.554)	467	2814 (3.554)	510	2865-2785	3.95 × 10 ⁻¹⁹	4.04 × 10 ⁻¹⁹
1297 (7.710)	16,980	1299 (7.698)	18,630	1345-1255	9.71 × 10 ⁻¹⁸	1.04 × 10 ⁻¹⁷

^a Determined from peaks in the absorbance spectrum.^b Determined from the calculated *k* values.



